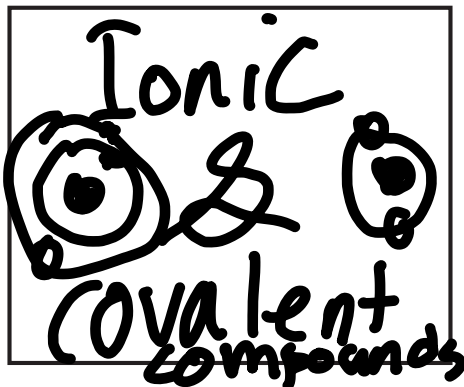
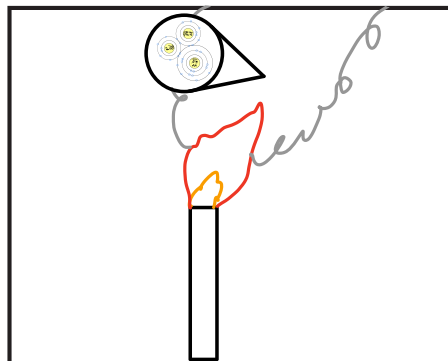


Storyboard Template

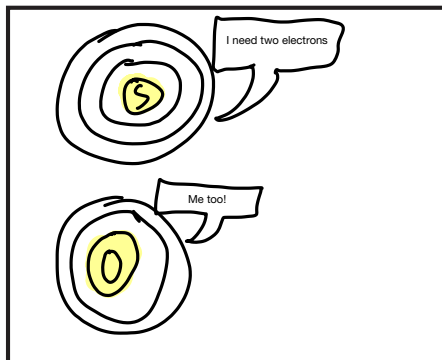
Title slide



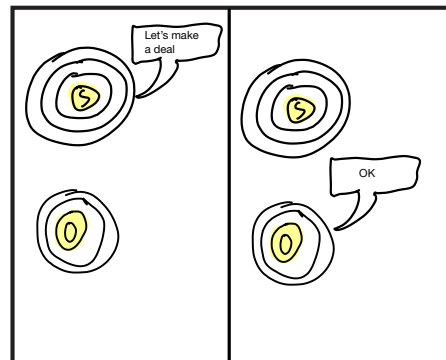
In this video, I am going to explain how the covalent compound Sulfur dioxide, forms and how the ionic compound calcium phosphide forms



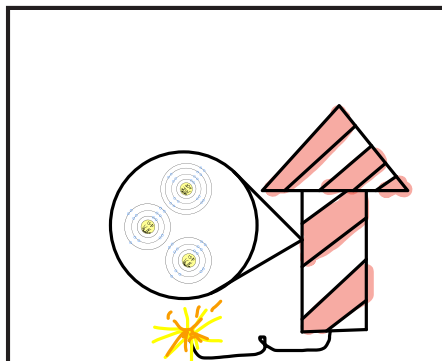
This forms sulfur dioxide, which is a covalent compound and a gas at room temperature. Sulfur dioxide is the gas responsible for the burnt smell from fire



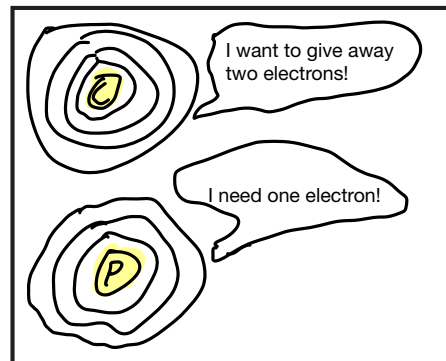
The covalent compound Sulfur dioxide forms because naturally occurring sulfur needs two electrons to complete their valence shell, and oxygen needs two electrons to complete its valence shell



The two oxygen atoms and the sulfur atom share their electrons and complete their valence shells



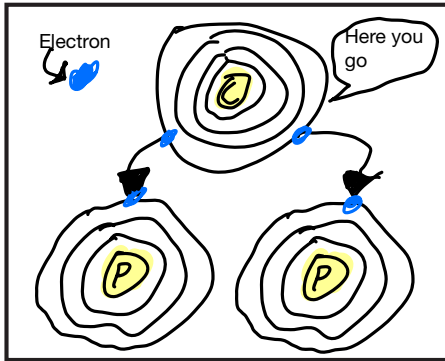
Calcium phosphide is an ionic compound commonly used in torpedos, fireworks, and flares.



Calcium phosphide consists of one calcium atom and two phosphorus atoms. Calcium needs to give away two electrons to complete its valence shell, and a phosphorus atom wants one electron

Name: Theryn

Project: Chemistry stories



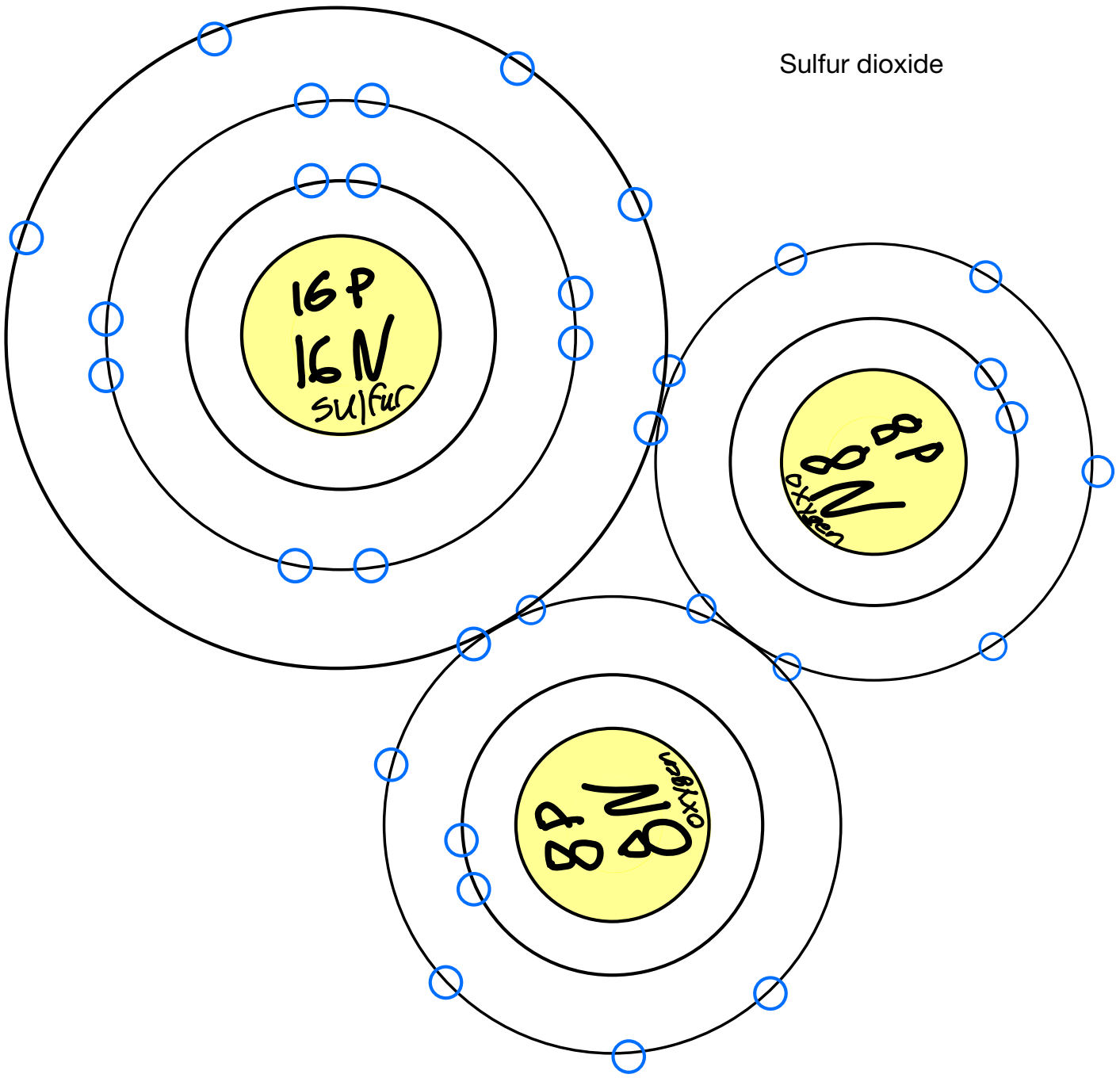
The calcium atom gives away one electron to each of the phosphorus atoms completing the calcium's and phosphorus's valence shells

the End

End scene



Sulfur dioxide



Calcium phosphide

