

Electronic Configuration PRACTICE QUIZ

For EACH of the following questions, on a SEPARATE piece of paper, write the:

- Full electronic configuration
- CORE electronic configuration
- Orbital Representation diagram (based from CORE electronic configuration)
- Lewis Dot Diagram

Although some of these will have QUITE lengthy answers for their Full Configurations, it is IMPERATIVE that you are clear on the order of filling of orbitals and how the numbers and letters change from $1s^2$ all the way to $7p^6$

If you choose not to write the full configurations for some of these practice questions, just be aware that you should expect at LEAST one or two LENGTHY questions that require you to list the FULL electron configuration, on the ACTUAL Quest.

1.	$_{14}\text{Si}$
2.	$_{33}\text{As}$
3.	$_{45}\text{Rh}$
4.	$_{73}\text{Ta}$
5.	$_{90}\text{Th}$
6.	$_{12}\text{Mg}^{+2}$
7.	$_{27}\text{Co}^{+2}$
8.	$_{27}\text{Co}^{+4}$
9.	$_{43}\text{Tc}^{+2}$
10.	$_{105}\text{Db}^{+5}$
11.	$_{108}\text{Hs}^{+2}$
12.	$_{15}\text{P}^{-3}$
13.	$_{52}\text{Te}^{-2}$
14.	$_{85}\text{At}^{-1}$
15.	$_{46}\text{Pd}^{+2}$

You can now practice with ANY of the 118 elements on the periodic table. You should be able to predict what the most likely ions (charges) would be by looking at the orbital representation diagrams.